

A sensitive spectrophotometer method for the determination of 4-aminoantipyrine as a Schiff base complicated with Cr (III), Ni (II) & Cu (II) transition elements and its preparation.

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Abstract

A sensitive spectrophotometer method for the determination of 4-aminoantipyrine in pure as well as in pharmaceutical dosage form is reported, based on the preparation of Schiff base as a chalet ligand by condensation reaction of 4-aminoantipyrine with Isatin to give (1-phenyl-2,3dimethyl-4Isatidine-5-oxopyrazole) followed by the preparation of three new complexes for this ligand with Cr(III),Ni(II),and Cu(II) in alcoholic media. These prepared complexes were characterized by C.H.N.,F.T.I.R.,electronic spectra,molar conductivity,magnetic sensitivity and measurements of metal ratio by flame atomic absorption technique.Beer's law was obeyed in the concentration ranges (5-400) $\mu\text{g.ml}^{-1}$, (7-300) $\mu\text{g.ml}^{-1}$ & (10-250) $\mu\text{g.ml}^{-1}$ with a detection limit(D.L) of (0.18) $\mu\text{g.ml}^{-1}$, (0.25) $\mu\text{g.ml}^{-1}$ & (0.33) $\mu\text{g.ml}^{-1}$ for Cr(III),Ni(II) & Cu(II) respectively. The method was successfully applied for the assay of pharmaceutical tables (Ampyrone).

الخلاصة

في هذا البحث تم استخدام طريقة طيفية لتقدير (4-aminoantipyrine) كمادة نقية وكذلك في المستحضر الصيدلاني، مبنية على تحضير قواعد شيف كليكاند مذبلي وذلك بتكاثف (1- فنيل-2،3، 4-ايساتيدين-5-او كسوباييرازول) ومن ثم تحضير ثلاث معقدات تناسقية لهذا الليكاند مع ايونات العناصر [الكروم (III) النيكل (II) والنحاس (II)] في وسط كحولي. وقد تم تشخيص هذه المعقدات بواسطة اجهزة التحليل الدقيق للعناصر (C.H.N.)، طيف الاشعة تحت الحمراء المعززة بتغيرات فورير، الاطياف الالكترونية، التوصيلية الكهربائية المولارية، الحساسية المغناطيسية وقياس نسبة الفلز بواسطة تقنية الامتصاص الذري اللهب. ان مدى التراكيز التي تطيع قانون بير هي (400-5) مايكرو غرام. مل⁻¹ (300-7) مايكرو غرام. مل⁻¹ و (250-10) مايكرو غرام. مل⁻¹ مع حدود كشف (0.18) مايكرو غرام. مل⁻¹، (0.25) مايكرو غرام. مل⁻¹ و (0.33) مايكرو غرام. مل⁻¹ على التوالي، ولقد تم تطبيق هذه الطريقة بنجاح لتقدير العقار الصيدلاني (Ampyrone).

1. Introduction

Pyrazolone is a five-membered lactam ring compound containing two nitrogen and ketone in the same molecule. Lactam structure is an active nucleus in pharmacological activity also lactam have big demand in artificial fibre industry, Pyrazolone is an active moiety as a pharmaceutical ingredient, especially in the class of nonsteroidal anti-inflammatory agents (NSAI) used in the treatment of arthritis and other musculoskeletal and joint disorders⁽¹⁾.

Antipyrine (also called phenazone) is a pyrazolone class analgesic agent in otic solutions in combination with other analgesic such as benzocaine, and phenylephrine. 4-aminoantipyrine (4-AAP) (4-Amino-2,3-dimethyl-1-phenyl-3-pyrazolin-5-one) is a metabolite of aminopyrine with an analgesic, anti-inflammatory, and antipyretic properties, it can be used as an intermediate for the synthesis of pharmaceuticals especially antipyretic and analgesic drugs. It is also used in the colorimetric determination of phenols⁽²⁾ (Fig.1).

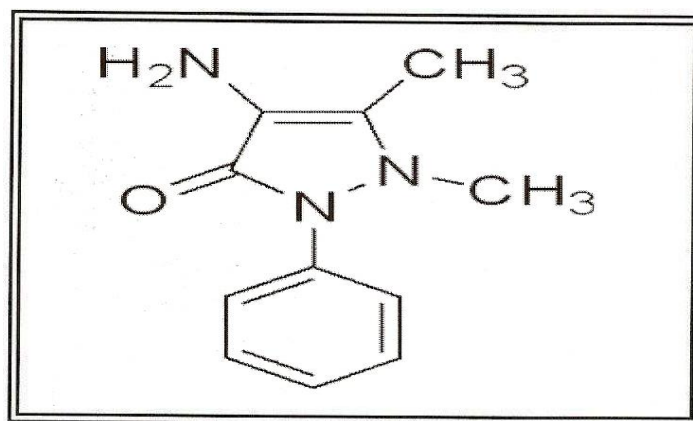


Fig (1).Chemical structure of (4-aminoantipyrine (4-AAP))

Analytical methods are useful in the determination of 4-AAP in pure and pharmaceutical dosage forms. Owing to the widespread use of UV-Visible Spectrophotometer in the routine analysis, it is important that good colorimetric methods are developed and that these are thoroughly validated^(3,4,5).

The proposed spectrophotometer method had sufficiently good accuracy, precision and permitted time-and money-saving assay of (Ampyrone).

2. Experimental

2-1 Apparatus and Reagents

Elemental microanalysis was carried out using C.H.N. elemental analyzer model Perkins-Elmer B240. Infra red spectra of the samples were recorded on F.T.I.R. 830 Shimadzu spectrophotometer using cesium iodide disc in the rang(200-400) cm^{-1} . The U.V.-Visible 160A-Ultra-Violet spectrophotometer in the rang 200-800nm. conductivity measurement were carried out by using SIMENS Digital Conductivity meter. Magnetic susceptibility measurements for the prepared complexes were obtained at room temperature using(magnetic susceptibility Balance) Johnson Matthey catalytic system division. Flame Atomic Absorption was obtained by using Shimadzu AA-670. Melting points of the prepared compounds were determined by using a calibrated meter F.B.S. Melting apparatus.

All chemicals used in this study were analat grad (Merk &Fluka).

2-2 Preparation of the Ligand (L)

An ethanolic solution of (20ml) Of 4-amino antipyrine (0.01mole, 2.3gm) was added to a solution of (0.01mole, 1.47gm) Isatin dissolved in (20ml) of ethanol. The mixture was refluxed on water bath and a red solid compound was separated as shown in scheme (1).

2-3 General Procedure

2-3-1 Preparation of calibration curve

A stock solution of the ligand was prepared by dissolving (0.05gm) in (50ml) volumetric flask, this solution was used to prepare the working standard solutions for different concentration between (5-400) $\mu\text{g}.\text{ml}^{-1}$ then add (100) $\mu\text{g}.\text{ml}^{-1}$ of $\text{CrCl}_3.6\text{H}_2\text{O}$, $\text{NiCl}_2.4\text{H}_2\text{O}$ & $\text{CuCl}_2.2\text{H}_2\text{O}$ separately, complete to the mark. The absorbance was measured at (532, 410 and 460) nm respectively.

2-3-2 Assay Procedure

A specific amount of powdered tables of (Ampyrone) equivalent to one table contain undergo condensing reaction with Isatin, react with $\text{CrCl}_3.6\text{H}_2\text{O}$, $\text{NiCl}_2.4\text{H}_2\text{O}$ & $\text{CuCl}_2.2\text{H}_2\text{O}$ respectively, then transfer into (25ml) volumetric flask and raise up to the mark. The drug concentration was determined from the calibration curve.

3. Result and Discussion

3-1 Physical and Analytical data

The physical and analytical data of the ligand and its metal complexes are given in table (1). Results obtained from elemental analysis are in satisfactory agreement with the calculated values.

3-2 Infra-red Spectra

The characteristic stretching vibration modes concerning ligand and its metal complexes were described in table(2). The IR spectrum of the ligand exhibited two distinct bands located at (1705 and 1630) cm^{-1} which are assigned to carbonyl group and isomethine group $\nu_{\text{C}=\text{N}}$ stretching respectively^(6,7).

A comparison of the IR spectrum of the free ligand and its metal complexes reveals that all complexes show the characteristic bands table (2). Coordination of the metal ion with nitrogen isomethine group $\nu_{\text{C}=\text{N}}$ is indicated by negative shift of $\nu_{\text{C}=\text{N}}$ by (15-20) cm^{-1} . Metal-nitrogen band $\nu_{\text{M} \rightarrow \text{N}}$ is further confirmed by the presence of the band around (470-480) cm^{-1} . As well as the negative shift of carbonyl group vibration in all the prepared complexes table (2). Support the participation of oxygen atom of carbonyl group to coordinate with the metal ion⁽⁸⁾.

New bands observed around the region (250-480) cm^{-1} were assigned to $\nu_{\text{M} \rightarrow \text{N}}$, $\nu_{\text{M} \rightarrow \text{O}}$, and $\nu_{\text{M} \rightarrow \text{Cl}}$ ⁽⁹⁾ modes table (2).

3-3 Electronic spectra, Molar conductance and Magnetic properties:

The UV spectrum of the ligand showed two bands at (245, and 341) nm that may be attributed to $\pi \rightarrow \pi^*$ and $n \rightarrow \pi^*$ respectively⁽¹⁰⁾ Fig(2).

3-3-1 Cr (III) complex

The bands at (13.135, 16.558 and 22.511) cm^{-1} in the Cr(III) complex assigned to the transitions (${}^4\text{A}_{2g}^{\text{F}} \rightarrow {}^4\text{T}_{2g}^{\text{F}}$), (${}^4\text{A}_{2g}^{\text{F}} \rightarrow {}^4\text{T}_{1g}^{\text{F}}$) and (${}^4\text{A}_{2g}^{\text{F}} \rightarrow {}^1\text{T}_{1g}^{\text{P}}$) respectively in an octahedral field^(1,2) Fig(3).

The value of magnetic moment ($\mu_{\text{eff}} = 3.68$ B.M.) showed that the complex expect to be high spin, the molar conductance of Cr (III) complex indicate the non electrolytic behavior table(3).

3-3-2 Ni (II) complex

The visible spectrum of the Ni (II) complex show multiple peak lie at (12.821, 17.458 and 27.155) cm^{-1} which are attributed to (${}^1\text{A}_{1g} \rightarrow {}^1\text{A}_{2g}$), (${}^1\text{A}_{1g} \rightarrow {}^1\text{B}_{1g}$) and (${}^1\text{A}_{1g} \rightarrow {}^1\text{E}_g$) respectively of square planar symmetry^(3,4) Fig(4), the value of magnetic moment ($\mu_{\text{eff}} = 0$ B.M.) showed that the

complex expected to be diamagnetic table(3),the molar conductance in DMF solvent indicate the electrolyte behavior in 1:1 ratio.

Table (1).Physical properties and C.H.N.M. elemental analysis of prepared compounds.

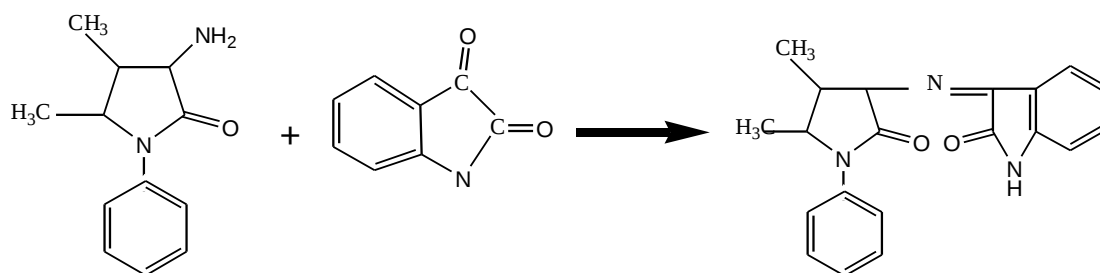
Symb.	Formula	m.p C°	% C.H.N.M. Calculated,(Found)			
			%C	%H	%N	%M
L	C ₁₉ H ₁₆ N ₄ O ₂	168-170	68.67 (68.33)	4.81 (4.25)	16.8 (16.34)	-----
C ₁	[Cr(C ₁₉ H ₁₆ N ₄ O ₂)Cl ₃]	233*	46.48 (46.05)	3.26 (3.11)	11.41 (10.85)	10.6 (9.87)
C ₂	[Ni(C ₁₉ H ₁₆ N ₄ O ₂)Cl]Cl	198*	49.38 (48.8)	3.46 (3.21)	12.12 (11.57)	12.7 (11.82)
C ₃	[Cu(C ₁₉ H ₁₆ N ₄ O ₂)Cl ₂ .(H ₂ O)]	220*	47.05 (46.25)	3.71 (3.51)	11.55 (10.92)	

*it means the decomposition of compound.

Table (2).Infrared spectra of the ligand and its Metal complexes.

Compound	$\nu_{C=N}$	$\nu_{C=O}$	ν_{M-N}	ν_{M-O}	ν_{M-Cl}
Ligand(L)	1630(s)	1705(s)	-----	-----	-----
[Cr(L)Cl ₃]	1610(m)	1690(m)	480	390	290(w)
[Ni(L)Cl]Cl	1615 (m)	1695(m)	470	390	250(m)
[Cu(L)Cl ₂ (H ₂ O)]	1610 (w)	1700 (m)	470	385	270(w)

*Where s=strong, w=weak, m=medium



Scheme (1): preparation of ligand (L).

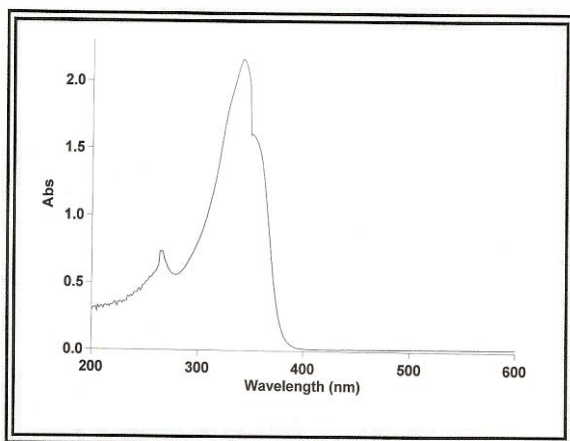


Fig (2): UV-Visible for ligand.

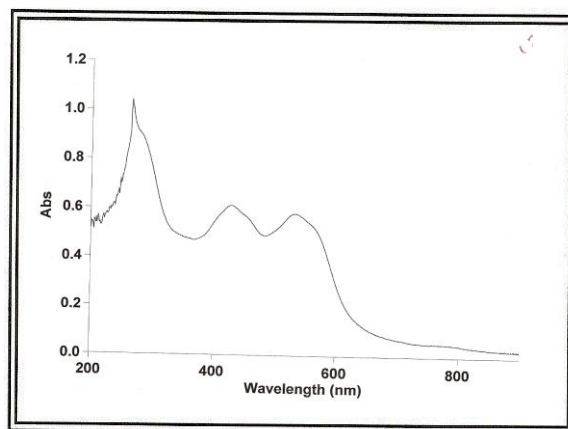


Fig (3): UV-Visible for C₁

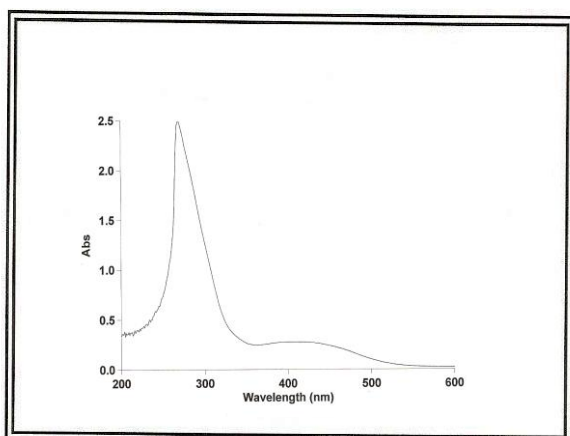


Fig (4): UV-Visible for C₂

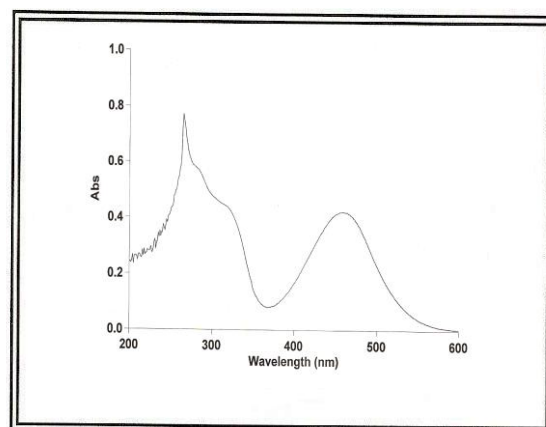


Fig (5): UV-Visible for C₃

Table (3).Electronic spectra data, Conductivity, Magnetic moment (B.M.)

Comp.	Electronic spectral peak (cm ⁻¹).	Tentative assignments	Λ_m Ohm ⁻¹ cm ² .mole ⁻¹	μ B.M.
L				
C ₁	13.135	$^4A_{2g}^F \rightarrow ^4T_{2g}^F$	3.2	3.68
	16.558	$^4A_{2g}^F \rightarrow ^4T_{1g}^F$		
	22.511	$^4A_{2g}^F \rightarrow ^1T_{1g}^P$		

C ₂	12.821	$^1A_{1g} \rightarrow ^1A_{2g}$	58.2	Diamagnetic
	17458	$^1A_{1g} \rightarrow ^1B_{1g}$		
	27155	$^1A_{1g} \rightarrow ^1E_g$		
C ₃	15.332	$^2E_g \rightarrow ^2T_{2g}$	15	1.68

3-3-3 Cu (II) complex:

The visible spectrum of Cu(II) complex show a broad band in the region(15.332)cm⁻¹ due to($^2E_g \rightarrow ^2T_{2g}$) of distorted octahedral complexes⁽⁵⁾ Fig(5).Conductivity value (Λ_m)in DMF showed that the complex was non-electrolyte, and the magnetic moment was(1.68)B.M. table(3).

According to the mentioned results, the suggested structure of the prepared complexes may be illustrated as follows:

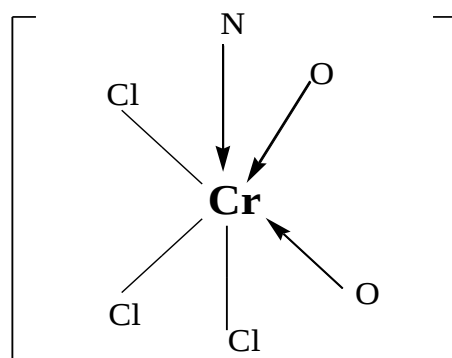


Fig (6): Facial (1-phenyl-2,3dimethyl-4Isatidine-5-oxo pyrazole tri chloro-chromium (III)).

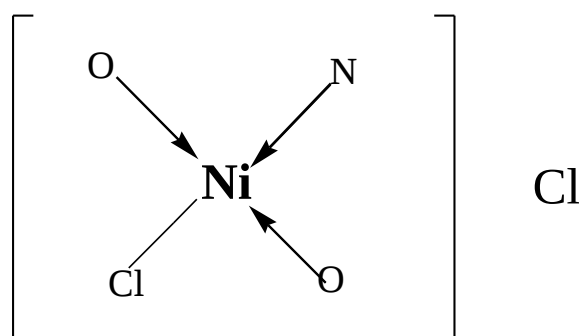


Fig (7):1-phenyl-2,3dimethyl-4Isatidine-5-oxopyrazole Nickel (II) chloride.

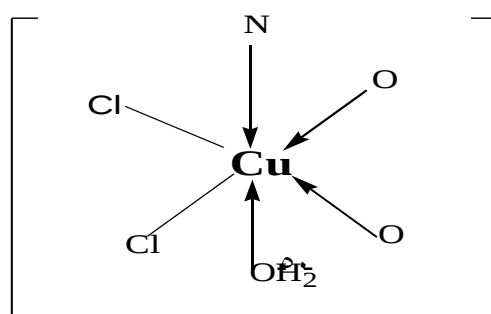


Fig (8):1-phenyl-2,3dimethyl-4Isatidine-5-oxopyrazole di chloro aqua
cupper (II).

3.4 Optical characteristics and validation of the method

The proposed procedure is validated by determining various optical parameters, which are listed in table(4).The linearity, slope and intercept have been calculated using the regression equation $y=ax+b$.The values of detection limit (D.L.), relative standard deviation (R.S.D.), confidence levels of con.and absor.,sandell's sensitivity(S),and molar absorptive coefficient(ϵ).

Table (4).Validation and sensitivity of the proposed method.

Parameters	$C_1[Cr(L)Cl_3]$	$C_2[Ni(L)Cl]Cl$	$C_3[Cu(L)Cl_2(H_2O)]$
Color	green	brownish	brownish
λ_{max}	532	410	460
Regre.Eq. $Y=ax+b$	$y=0.0017x+0.02$	$y=0.0054x+0.087$	$y=0.0057x+0.06$
Corr.coef.(r)	0.9999	0.9994	0.9997
Tabulated t-test two tailed (n-2) 95% C.I.	3.182	2.571	2.776
Conf.Limit for slop $b \pm ts_b$	0.0017 ± 0.00016	0.0054 ± 0.00061	0.0057 ± 0.0019
Conf.Limit for intercept $a \pm ts_a$	0.02 ± 0.096	0.087 ± 0.24	0.06 ± 0.36
D.L. $\mu g.ml^{-1}$ n=14	0.18	0.25	0.33
Erel. %	0.075	0.05	0.1
RSD%	1.15	2.02	1.18
Conf.Limit conc. $\mu g.ml^{-1}$ 95%C.I.	399.7 ± 0.94	199.9 ± 0.9	99.9 ± 1.2
Conf.Limit abs.	0.699 ± 0.0017	1.08 ± 0.004	0.557 ± 0.002

95%C.I.			
S $\mu\text{g.cm}^{-2}$	0.1	0.179	0.16
ϵ L.mol ⁻¹ .cm ⁻¹	2.8*10 ³	1.13*10 ³	1.19*10 ³

4. Applicability of the method

The applicability of the suggested method for the assay of (Ampyrone) was examined. The results of the assay of tables were given in table (5) for all complexes.

Table (5).Determination of 4-AAP in pharmaceutical preparation.

Commercial formulation analyzed	Label claim(mg)	C ₁ [Cr(L)Cl ₃]	C ₂ [Ni(L)Cl]Cl	C ₃ [Cu(L)Cl ₂ (H ₂ O)]
Ampyrone	25	24.937	24.975	24.9

Table (6).Recovery results of the proposed method.

complexes	C ₁ [Cr(L)Cl ₃]	C ₂ [Ni(L)Cl]Cl	C ₃ [Cu(L)Cl ₂ (H ₂ O)]
Recovery %	99.75	99.9	99.6

5. Interferences

To study the potential interference from the commonly used excipients and other additives such as starch, talc, magnesium stearate, and glucose, under the experimental conditions employed in this paper.

Recovery studies were carried out and the results of the recovery analysis, table (6) showed no interference with the assay.

6. Conclusion

The objective of the investigation reported in this paper was to evaluate a spectrophotometer method for the determination of 4-aminoantipyrine in pharmaceutical dosage form, based on the synthesis of the new Schiff base (1-phenyl-2,3-dimethyl-4-isatidine-5-oxopyrazole) with some metal complexes.

This method was found to be reliable, accurate, and more sensitive than most of the analytical methods reported.

References

- 1-A. Burcinova, and M. Tichy, J.Chromatogr., 25,455,420-4(1988).
- 2-M.Q. Al-Abchi, A.M. Al-Delami, and S. Al-Najafi, Analyst, 113(11), 1661-3(1988).
- 3-V.V. Zanin, E.V. Smirnov, and N.N. Papkian, Farmatsiia, 28(6), 33-7 (1988).
- 4-Wikipedia, the free encyclopedia (2007), Copy right by Wikimedia Foundation, Inc., United State, on line.
- 5-Sigma-aldrich (2007), by Sigma-Aldrich Co., on line.
- 6-M.S. Patial, and J.T. Shah, J. Indian Chem. Soc., 944, 58(1981).
- 7-A.B.P. Lever. "Inorganic Electronic Spectroscopy" Amsterdam-London, (1968).
- 8-B. Singh, and H. Misra, J.Indian Chem.Soc., LX III, 692(1988).
- 9-P. Venkates, B. War, N. Ramaro, and V. Tyaga, Indian J. of Chem., 25A, 482-484(1986).
- 10-T. Sarojini, A. Ramch, and C. Raiah, Indian J. of Chem., 35A:940-945 (1996).
- 11-C.B. Mahto, Indian J. of Chem.Soc., 58,935(1981).
- 12-V.J. Baker, D.V. Kasnis, J. Indian Chem. Soc., 58, 790(1978).
- 13-L. Mishra, and K. Mustafa, Indian J.of Chem., 32A, 917(1993).
- 14-K.Nakamoto, "Infrared Spectra of Inorganic and coordination compounds", 4th ed. Wiley, Interscience, New York (1986).
- 15-K.Brugger "Coordination chemistry: Experimental methods" London Butterworths, England (1973).